

TABLE 5. Redox potential of electron donors and electron acceptors involved in ETP

Redox compound	E'_0 (mV)	References
SO ₄ ²⁻ /HSO ₃ ⁻	-516	- ^b
CO ₂ /formate ^a	-432	- ^b
H ⁺ /H ₂	-414	- ^c
S ₂ O ₃ ²⁻ /HS ⁻ + HSO ₃ ⁻	-402	- ^b
Flavodoxin ox/red (E'_{01})	-371 ^d	397
Ferredoxin ox/red (E'_{01})	-398 ^e	157
NAD/NADH	-320	91
Cytochrome c ₃ ox/red	-290	728
CO ₂ /acetate ⁻	-290	- ^b
S ⁰ /HS ⁻	-270	- ^b
CO ₂ /CH ₄	-244	- ^b
FAD/FADH ₂	-220	- ^f
Acetaldehyde/ethanol	-197	- ^b
Pyruvate ⁻ /lactate ⁻	-190	- ^b
FMN/FMNH ₂	-190	- ^g
Dihydroxyacetone phosphate/glycerol-phosphate	-190	94
HSO ₃ ⁻ /S ₃ O ₆ ²⁻	-173	- ^b
Oxaloacetate ²⁻ /malate ²⁻	-172	- ^b
Flavodoxin ox/red (E'_{02})	-115 ^d	397
HSO ₃ ⁻ /HS ⁻	-116	- ^b
Menaquinone ox/red (MK)	-74	569, 685
APS/AMP + HSO ₃ ⁻	-60	594
Rubredoxin ox/red	-57	155
Acrylyl CoA/propionyl CoA	-15	232
Glycine/acetate ⁻ + NH ₄ ⁺	-10	- ^b
2-Demethylvitamin K ₂ ox/red	+25	569, 250a
S ₄ O ₆ ²⁻ /S ₂ O ₃ ²⁻	+24	- ^b
Fumarate/succinate	+33	- ^b
Ubiquinone ox/red	+113	569
S ₃ O ₆ ²⁻ /S ₂ O ₃ ²⁻ + HSO ₃ ⁻	+225	- ^b
NO ₂ ⁻ /NO	+350	- ^b
NO ₃ ⁻ /NO ₂ ⁻	+433	- ^b
Fe ³⁺ /Fe ²⁺	+772	- ^b
O ₂ /H ₂ O	+818	- ^b
NO/N ₂ O	+1175	- ^b
N ₂ O/N ₂	+1355	- ^b

^a CO₂ rather than HCO₃⁻ has been shown to be the active species of "CO₂" utilized or formed by formate dehydrogenases (336, 645, 646b).

^b Calculated from ΔG° for redox compound reduction with H₂ ($\Delta G^{\circ} = -n \cdot F \cdot \Delta E'_0$, $\Delta E'_0 = E'_0$ (redox compound) - E'_0 (H⁺/H₂); n = number of electrons transferred in the reaction; F [Faraday constant] = 23060.9 cal/V equivalent) (110). The ΔG° values were calculated from the ΔG_f° values given in Table 15 (CO₂, CH₄, H₂, N₂, NO, and N₂O in the gaseous state, all other compounds in aqueous solution).

^c At 25°C.

^d *Peptostreptococcus elsdenii*; for E'_0 of clostridial flavodoxins see (394).

^e *Clostridium pasteurianum* ($E'_{02} = -367$ mV).

^f The redox potential of flavin enzymes may differ by as much as 200 mV from the values of the free coenzymes.